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Terms in this set (19)

Ion. An atom that  
gained or lost electrons  
and has a charge.

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Compounds And  
Answers  
Cation-Lost electron-  
Positive charge-Metals.  
Anion-Gained electrons-  
Negative charge-  
Nonmetals. Properties.

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The energy required to separate 1 mol of ions in an ionic compound is referred to as lattice energy. Lattice energy is directly related to the size of the ions that are bonded. Smaller

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ions form compounds with more closely spaced ionic charges and require more energy to separate. Smaller, harder to get between them.

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Compounds and Metals  
Section 77.1.1 Figure  
7.1 As carbon dioxide  
dis-solves in ocean  
water, carbonate ions

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are produced. Coral polyps capture these car-bonate ions, producing crystals of calcium carbonate, which they secrete as an exo-skeleton. Over time, the coral reef forms. A coral reef is a complex habitat that

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Chapter 7 Review:  
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response in the blank,  
at the left, that best  
completes the  
following: \_\_\_\_\_ 1. \_\_\_\_\_  
is the force that holds  
2, or more, atoms  
together. \_\_\_\_\_ 2. \_\_\_\_\_  
electrons are those  
electrons that are in  
the outer energy level.  
\_\_\_\_\_ 3.

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Valence Electron. Octet  
Rule. To find the  
number of valence

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electrons.... Atoms of the metallic elements tend to.... An electron that is in the highest occupied energy level of an.... In forming compounds, atoms tend to achieve the electron confi....

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## CHAPTER 7 SOLUTIONS

MANUAL 25. calcium  
and chlorate  $\text{Ca}(\text{ClO}_3)_2$

26. aluminum and  
carbonate  $\text{Al}_2(\text{CO}_3)_3$

27. Challenge Write the  
formula for an ionic  
compound formed by  
ions from a group 2  
element and  
polyatomic ions  
composed of only of  
carbon and oxygen.  
The polyatomic ion is  
carbonate ( $\text{CO}_3^{2-}$ ).

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Ionic Bonds. Displaying  
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(Practice Quiz) (with  
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## **Quia - Chapter 7 Ionic Compound**

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AND CHEMICAL  
COMPOUNDS 203  
SECTION 7-1 O

OBJECTIVES Explain the

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significance of a  
chemical formula.

Determine the formula of an ionic compound formed between two given ions. Name an ionic compound given its formula. Using prefixes, name a binary molecular compound from its formula. Write the formula of a binary molecular compound given its name. C

## **CHAPTER 7 Chemical Formulas and**



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## Compounds And Worksheet

### **Chemical Compounds**

A compound that contains ions and is held together by ionic bonds is called an ionic compound. The periodic table can help us recognize many of the compounds that are ionic: When a metal is combined with one or more nonmetals, the compound is usually ionic.

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ionic.

## 2.7 Molecular and Ionic Compounds | Chemistry

Naming Ions. The name of a monatomic cation is simply the name of the element followed by the word ion. Thus,  $\text{Na}^+$  is the sodium ion,  $\text{Al}^{3+}$  is the aluminum ion,  $\text{Ca}^{2+}$  is the calcium ion, and so forth.. We have seen that some elements lose different

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numbers of electrons,  
producing ions of  
different charges  
(Figure 3.3).

## **5.7: Naming Ionic Compounds - Chemistry LibreTexts**

Ionic Bonds - Chapter 7

1. Ionic Bonding and  
Ionic Compounds Or  
How I Learned to Love  
Electrostatic Forces 2.  
Ionic Bonds The  
periodic table is based  
upon the properties of

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the elements. After the discovery of the electron, scientists found that all elements in the same group have the same number of valence electrons.

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25. State the number  
of electrons lost or  
gained in forming each

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of these ions. Name  
the ions and tell  
whether it is an anion  
or a cation.

**05 CTR ch07 7/9/04  
3:27 PM Page 155  
IONS 7**

Yay, this practice test  
will determine if you're  
ready to take the real  
thing. Questions are  
taken right out of the  
book/packet and  
review lotsa vocab.

**Chapter 7 Chem:**  
*Page 23/25*

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ions attract each other, forming ionic bonds.

The resulting compounds are called ionic compounds and are the primary subject of this chapter. The second way for an atom to obtain an octet of electrons is by sharing electrons with another atom. These shared electrons simultaneously occupy the outermost shell of



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more than one atom.

The bond

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